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NASA CASE NO. LAR-13732-1

PRINT FIG. 1

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{NASA-Case-LAR-13732-1} PROCESS FOR  
DEVELOPING CRYSTALLINITY IN LINEAR AROMATIC  
POLYIMIDES Patent Application (NASA) 29 p  
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G3/27 0083857

PROCESS FOR DEVELOPING CRYSTALLINITY  
IN LINEAR AROMATIC POLYIMIDES

AWARDS ABSTRACT

NASA Case No. LAR 13732-1

This invention relates to the chemical dehydration of polyamide-acids to form polyimides which are semi-crystalline, and which have melting temperatures near the glass transition temperatures of the amorphous forms of the polyimides. Upon melting of these semi-crystalline polyimides, the crystalline forms do not reform. Flow is thereby enhanced as melting occurs. Moreover, the semi-crystalline polyimides are prepared without degradation of the molecular weight of the polyimides.

The process of the present invention includes first treating a polyamide-acid (such as LARC-TPI polyamide-acid) in an amide-containing solvent (such as N-methylpyrrolidone) with an aprotic organic base (such as triethylamine), followed by dehydrating with an organic dehydrating agent (such as acetic anhydride).

The novelty of this invention resides in the particular sequence of procedural steps which constitute a chemical dehydration process for the preparation of semi-crystalline polyimides from the polyamide-acid precursors thereof.

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Initial Evaluator:

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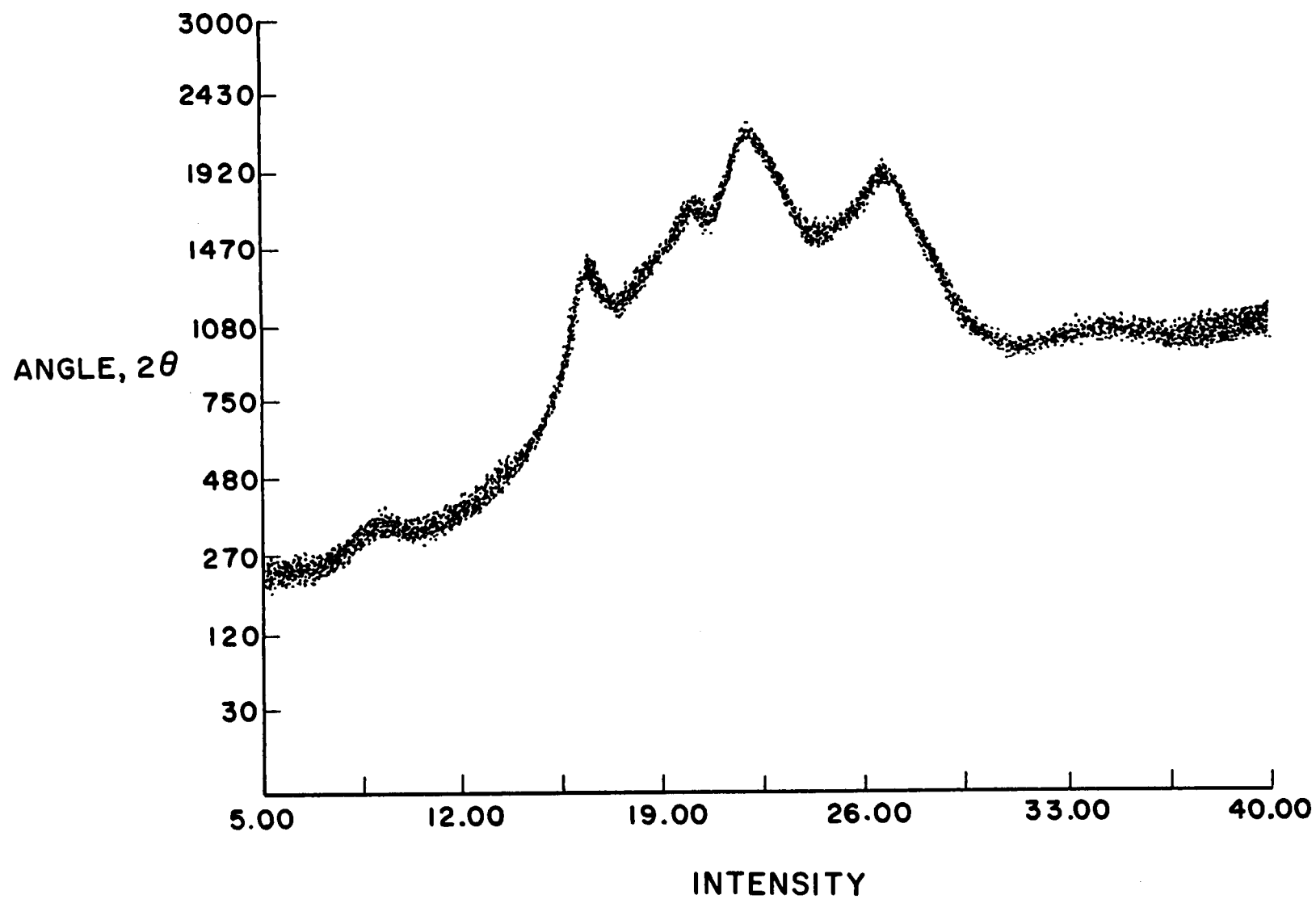


FIG. 2

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SHEET 2 OF 10  
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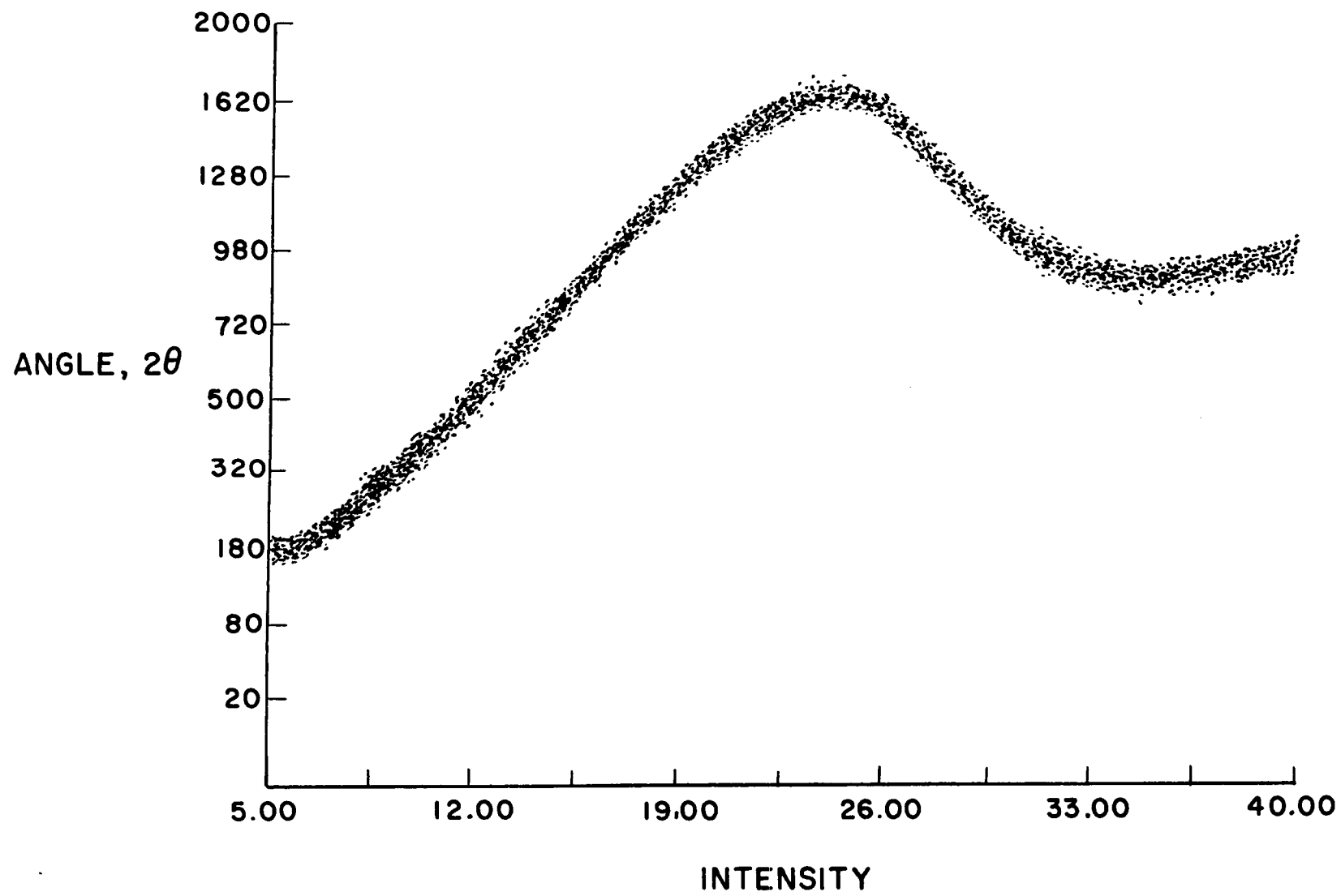


FIG. 1

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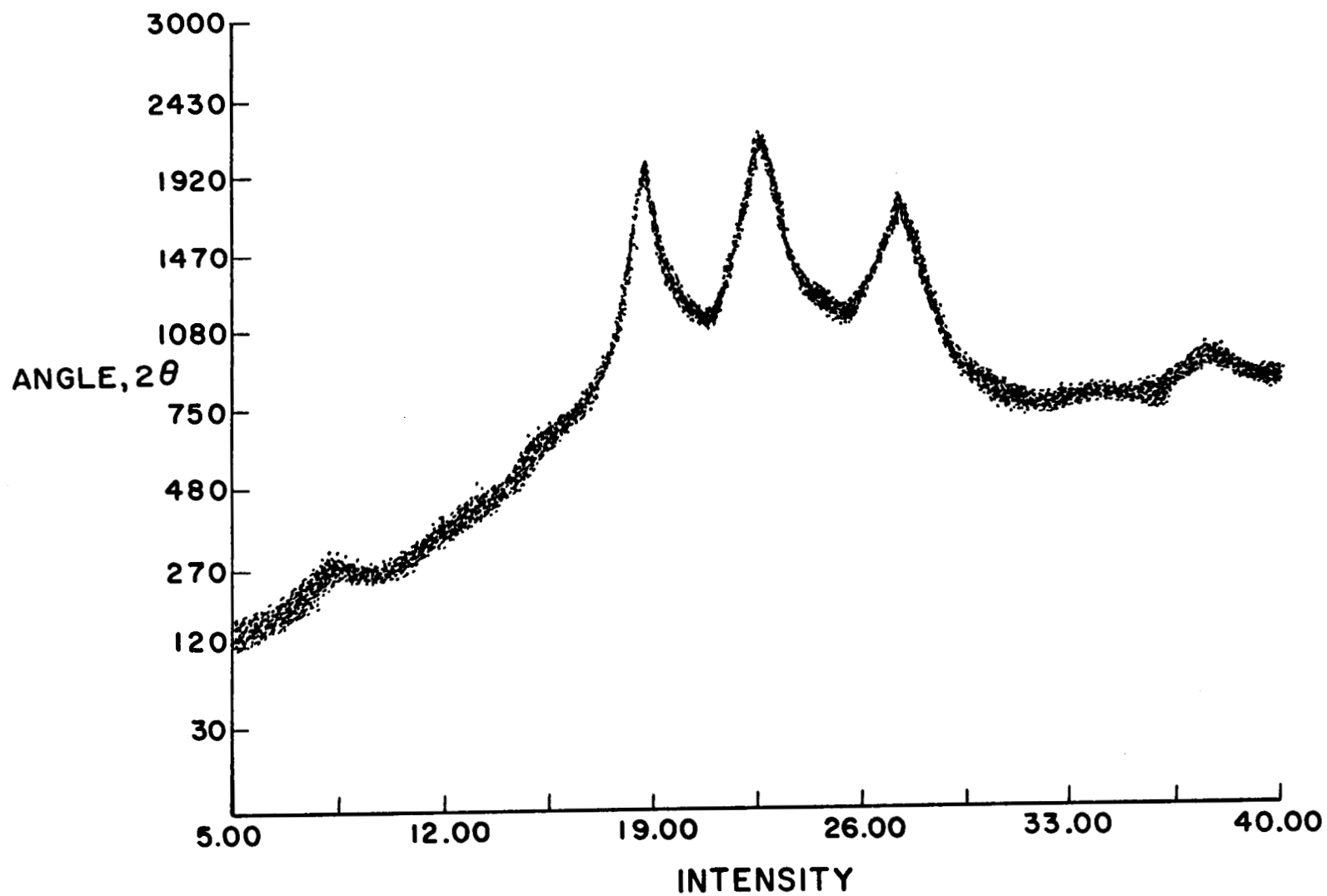


FIG. 3

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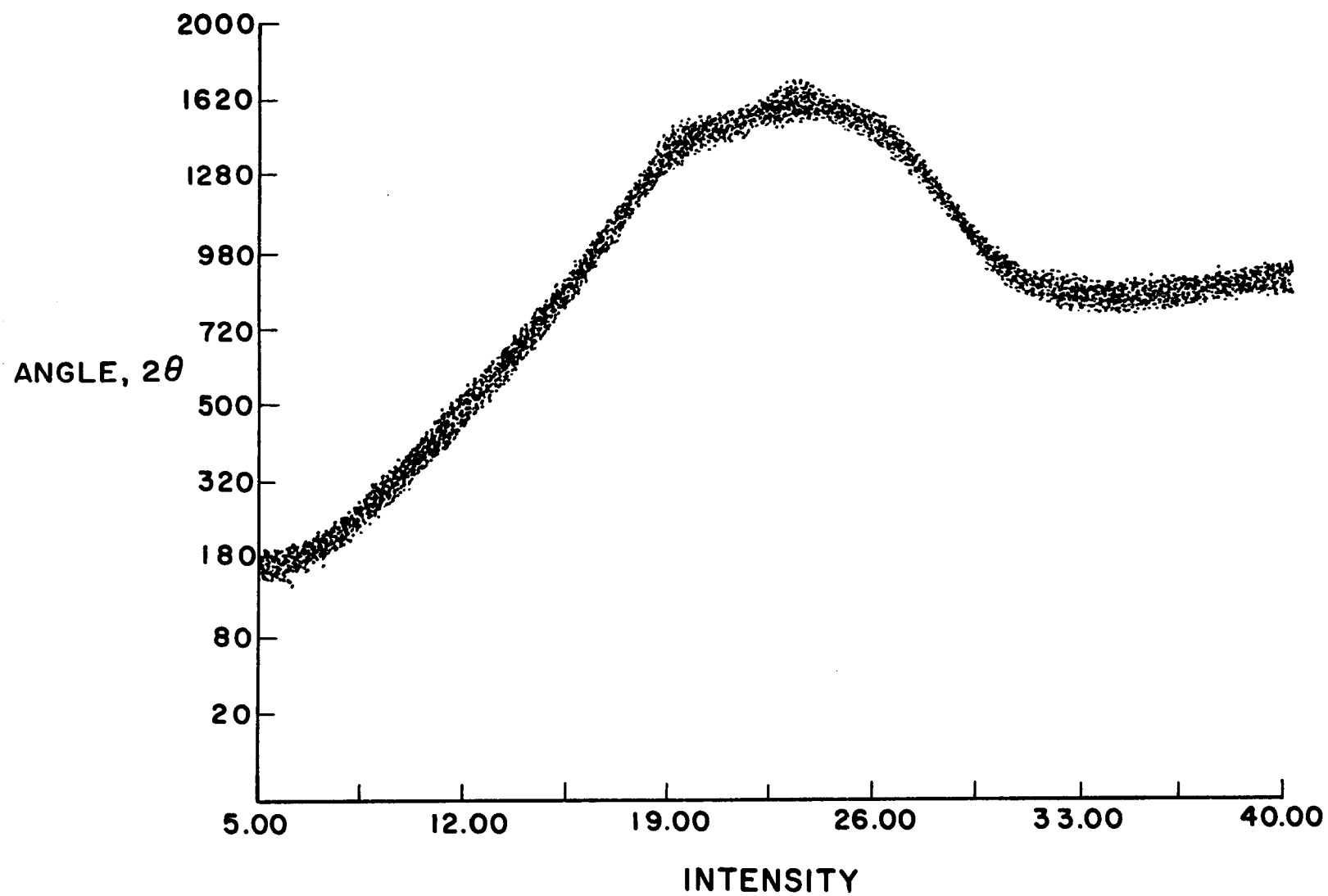


FIG. 4

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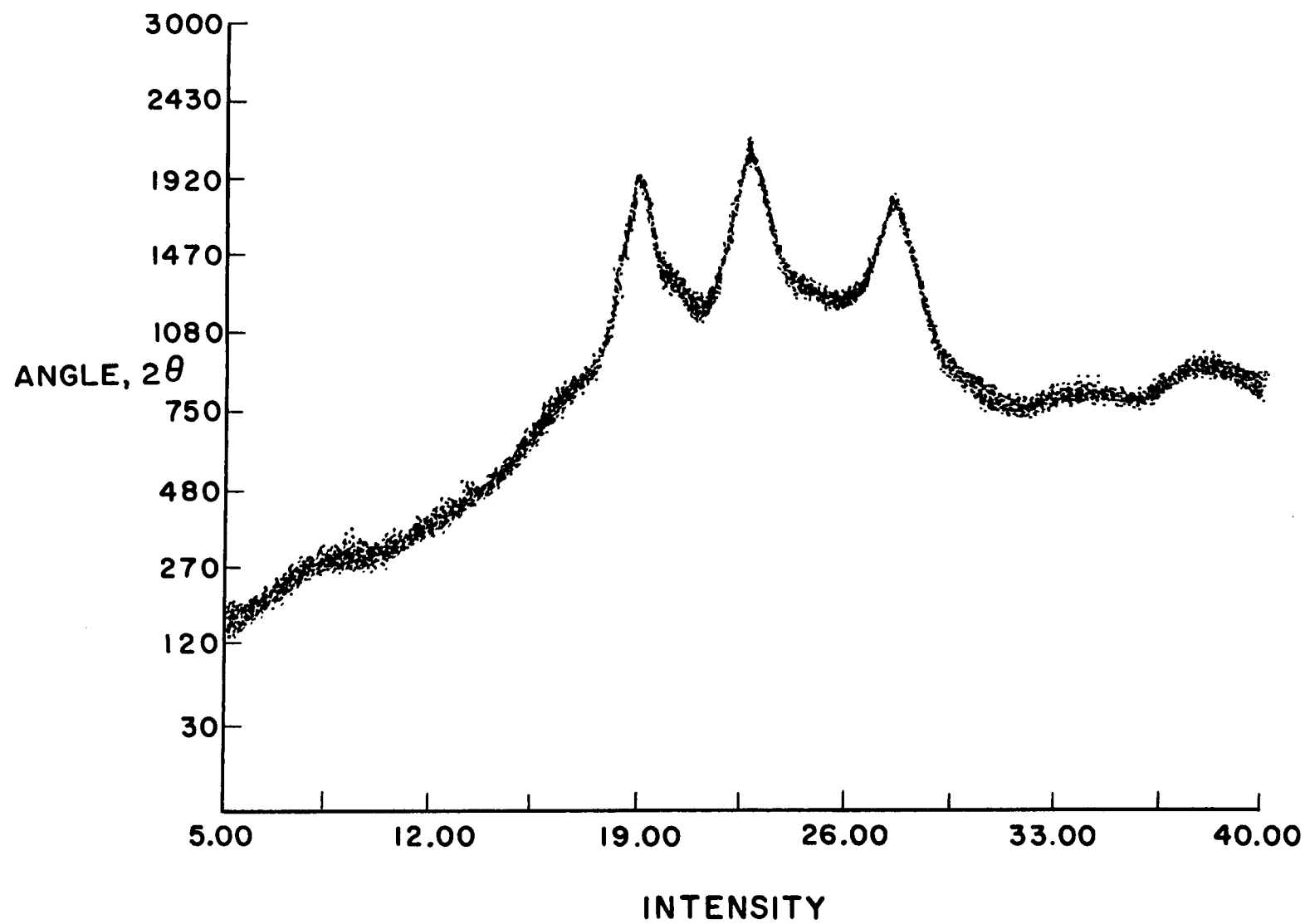


FIG. 5

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SHEET 5 OF 10  
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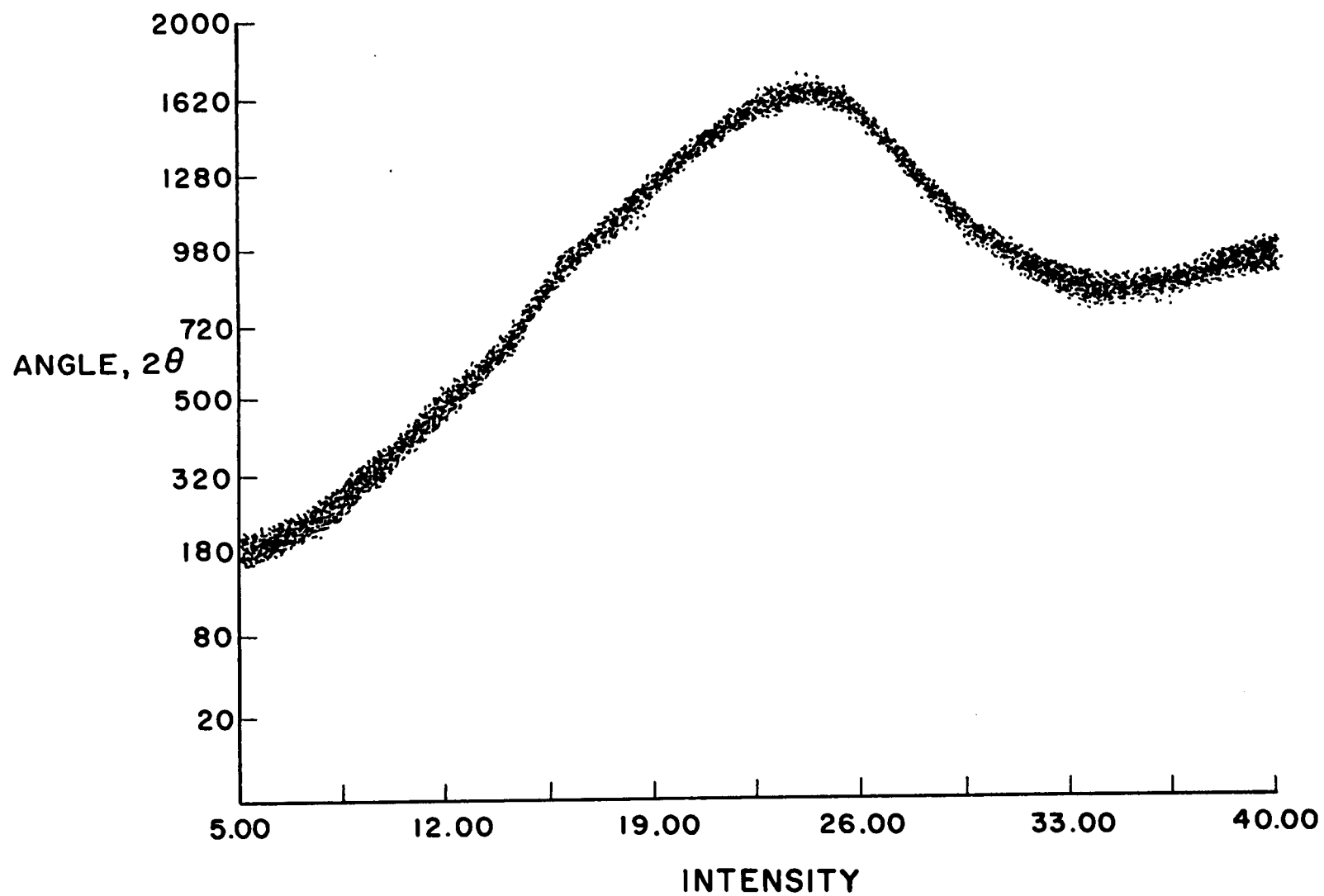


FIG. 6

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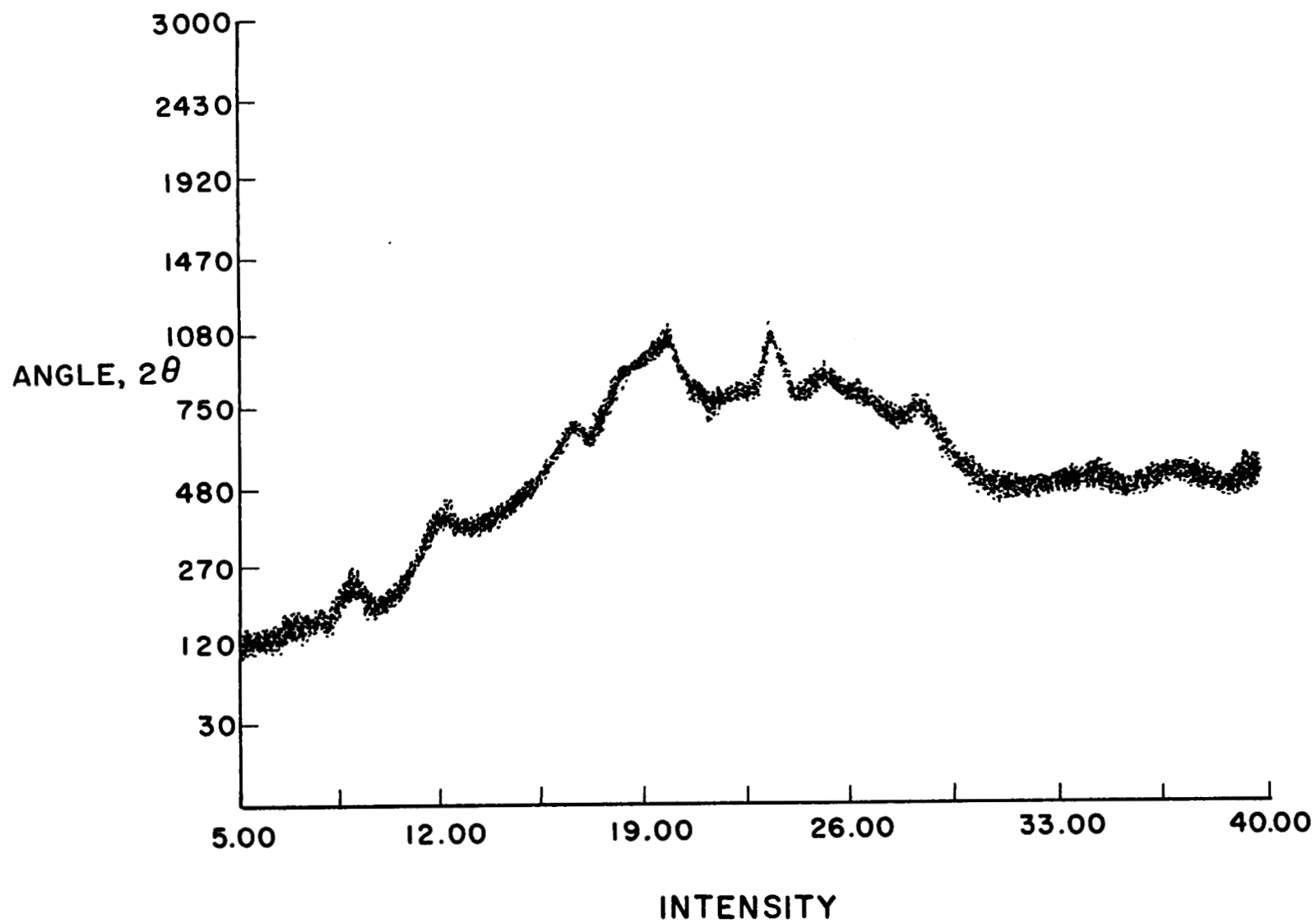


FIG. 7

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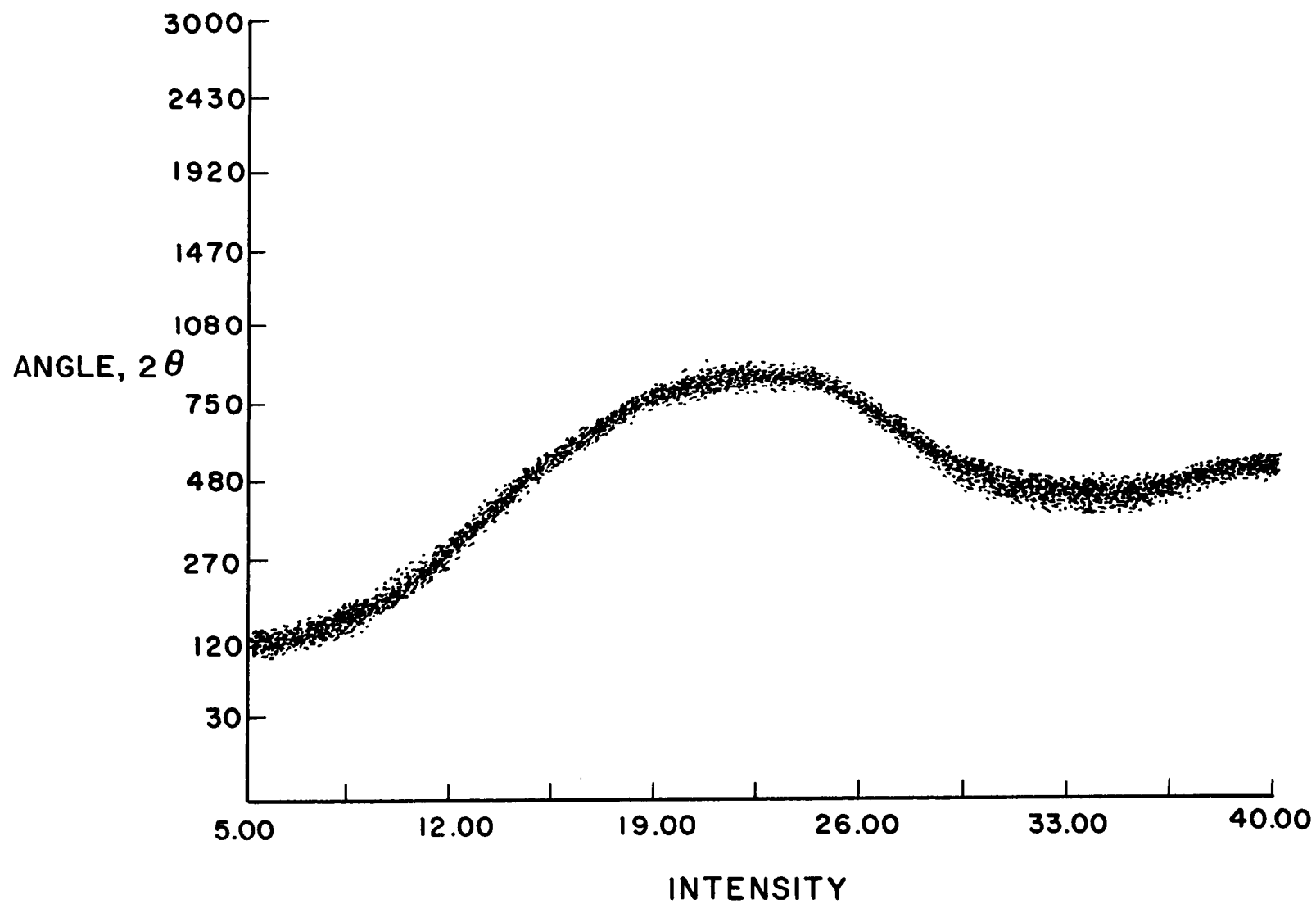


FIG. 8

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SHEET 8 OF 10

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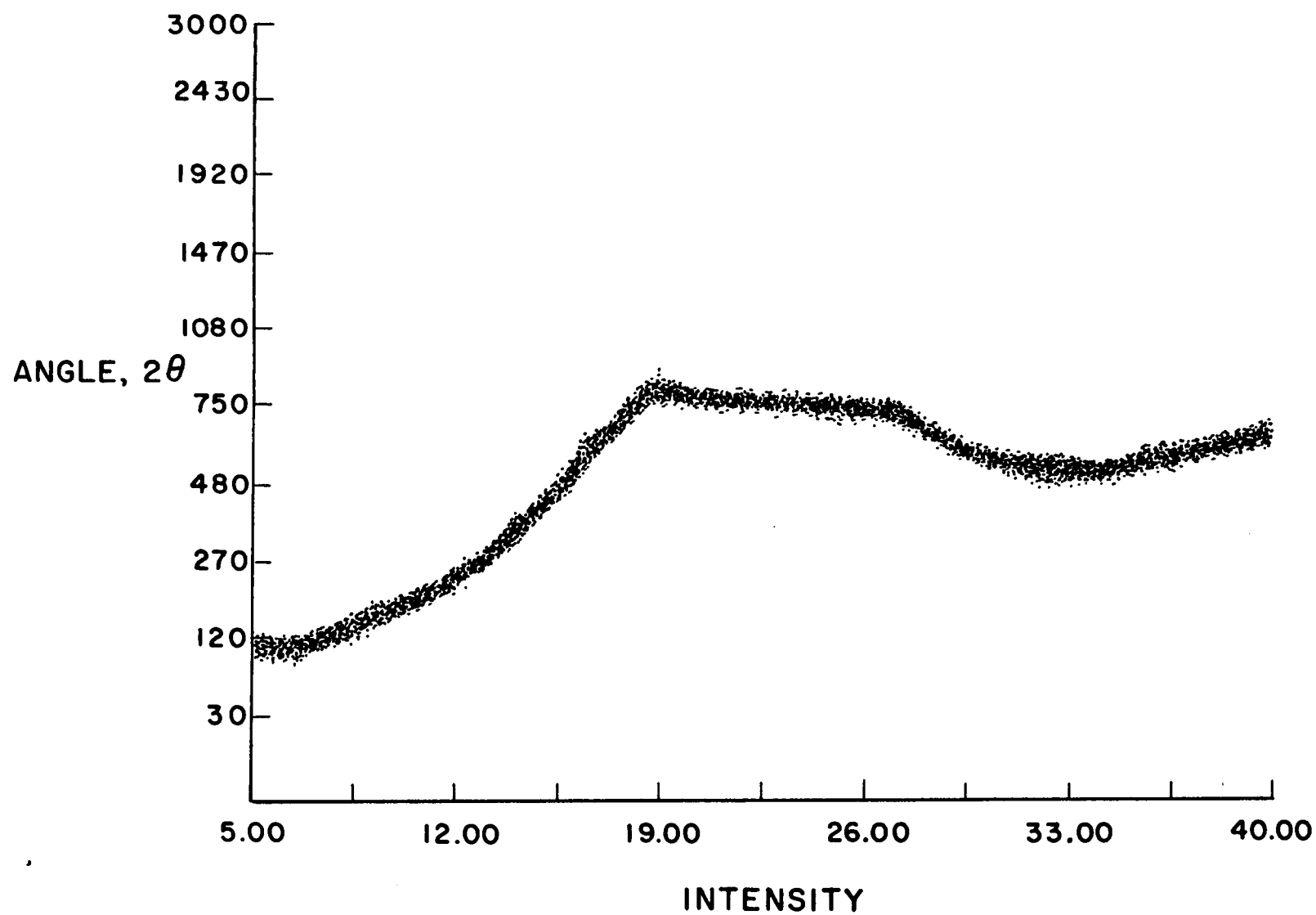


FIG.9

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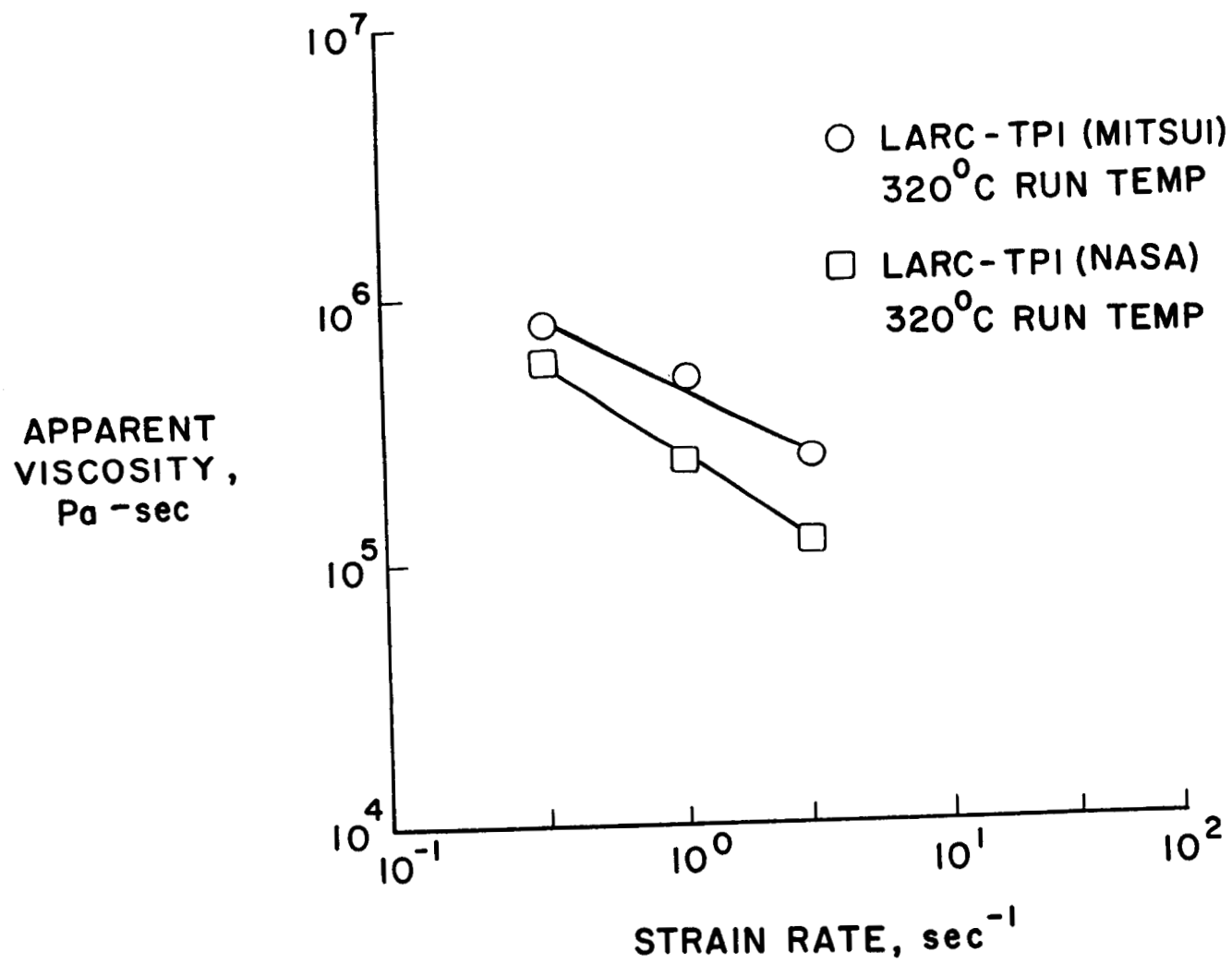


FIG. 10

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SHEET 10 OF 10  
INVENTOR(S): TERRY L. ST. CLAIR

STAR

PROCESS FOR DEVELOPING CRYSTALLINITY IN LINEAR AROMATIC  
POLYIMIDES

5    Origin of the Invention

          The invention described herein was made by an employee  
of the United States Government and may be manufactured and  
used by or for the Government for governmental purposes  
10    without the payment of any royalties thereon or therefor.

Background of the Invention

Field of the Invention

15           This invention relates generally to linear aromatic  
polyimides. It relates particularly to preparing  
semi-crystalline polyimides without degrading the molecular  
weight thereof.

20    Prior Art

          Linear aromatic polyimides are well known to have  
exceptional thermal and oxidative stabilities. However,  
25    they are very difficult to process into useful forms,  
because of their poor softening/flow properties. Most of  
these polymers have to be processed at temperatures  
approaching 400°C and at pressures of 500 to 10,000 psi. At  
these high temperatures and pressures a sintering of finely  
30    divided polymer particles occurs. In order for linear  
aromatic polyimides to be useful as adhesives, molding  
powders and matrix resins for various applications, the  
polymers should exhibit flow at lower temperatures and  
pressures.

35           Accordingly, a primary object of the present invention  
is to provide a process for converting the polyamide-acid  
precursors of polyimides to polyimides which exhibit  
significant levels of crystallinity, so that these  
crystalline forms can then be melted at relatively moderate

temperatures to afford adequate flow for molding and consolidation operations.

5 Another object of the present invention is to provide a method for controlling the level of crystallinity that develops in polyimides as they are chemically cyclodehydrated from their corresponding polyamide-acid precursors.

10 A further object of the present invention is to provide a process for chemically converting polyamide-acids to the corresponding semi-crystalline polyimides without degrading the molecular weight of the polyimides.

#### Summary of the Invention

15 By the present invention, linear aromatic polyimides are produced which exhibit a form of metastable crystallinity that results in a melt-flow behavior that makes these materials attractive as adhesives, molding powders and matrix resins for many potential applications.

20 According to the present invention, the foregoing and additional objects are attained by providing a process which involves first dissolving a polyamide-acid in an amide solvent, or a mixture of an ether and an amide solvent, at low percent solids, followed by treating this solution with an aprotic organic base, such as triethylamine or pyridine. Finally, this solution is treated with an organic dehydrating agent.

30 It is important to observe that with regard to the dissolving of the polyamide-acid in the solvent, at least some amide solvent must be employed to afford crystallinity.

35 Moreover, it is important to observe that with regard to the treatment of the polyamide-acid in dilute solution with an aprotic organic base, the solution must be stirred or allowed to be in the presence of the aprotic organic base prior to the addition of the dehydrating agent, or crystallinity will not develop.

40 Finally, it is important to observe that with regard to the treatment of the dilute polyamide-acid/aprotic organic base solution with the dehydrating agent, the solubility of



the resulting polyimide in the reaction mixture affects the resulting degree of crystallinity. Also, the time that the dehydrated polymer is in contact with the final solution affects the degree of crystallinity.

#### Brief Description of the Drawings

For a more complete understanding of the present invention, including its objects and benefits, reference should be made to the Detailed Description of the Preferred Embodiments, which is set forth below. This Detailed Description should be read together with the accompanying drawings, wherein:

FIG. 1 is a representation of the X-ray diffraction pattern for the amorphous form of a linear aromatic polyimide designated LARC-TPI;

FIG. 2 is a representation of the X-ray diffraction pattern for a commercially-available semi-crystalline form of the linear aromatic polyimide LARC-TPI;

FIG. 3 is a representation of the X-ray diffraction pattern for the linear aromatic polyimide LARC-TPI as prepared according to a first embodiment of the present invention;

FIG. 4 is a representation of the X-ray diffraction pattern for the linear aromatic polyimide LARC-TPI as prepared by a first process which is similar to, but not the same as, the present invention;

FIG. 5 is a representation of the X-ray diffraction pattern for LARC-TPI which is prepared according to a second embodiment of the present invention;

FIG. 6 is a representation of the X-ray diffraction pattern for LARC-TPI which is prepared by a second process which is similar to, but not the same as, the present invention;

FIG. 7 is a representation of the X-ray diffraction pattern for the linear aromatic polyimide polyimidesulfone (PISO2) as prepared according to another embodiment of the present invention;

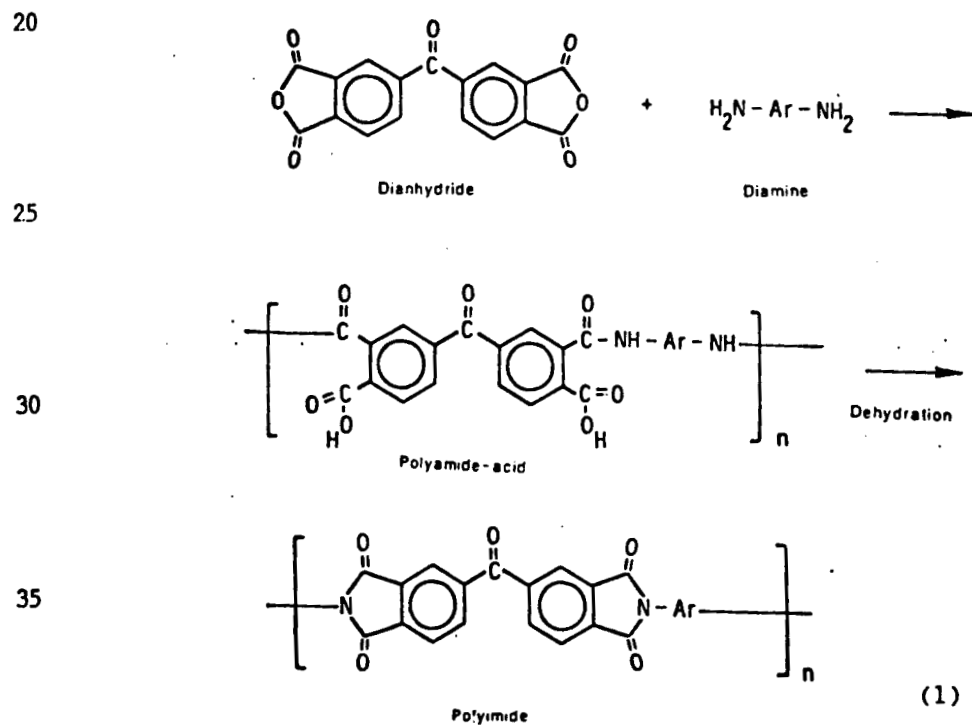
FIG. 8 is a representation of the X-ray diffraction pattern for PISO2 which was imidized by heating;

FIG. 9 is a representation of the X-ray diffraction pattern for the linear aromatic polyimide BTDA-ODA as prepared according to another embodiment of the present invention; and

FIG. 10 is a plot of apparent viscosity as a function of strain rate for LARC-TPI which is prepared according to the present invention and for semi-crystalline LARC-TPI which is available commercially.

#### Detailed Description of the Preferred Embodiments

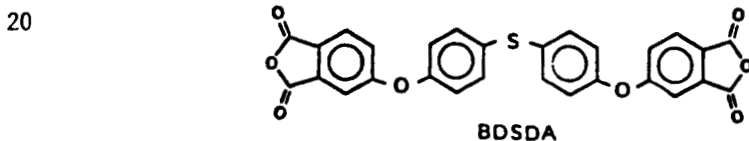
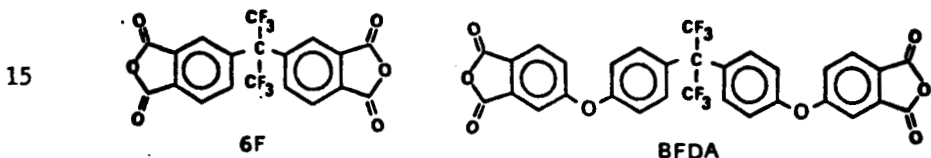
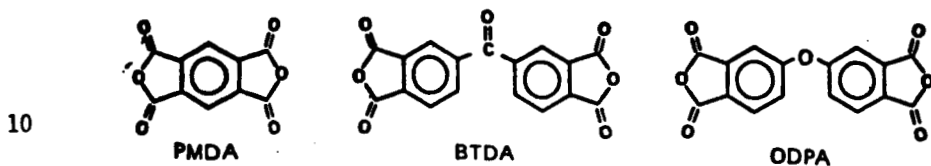
The scheme for preparing linear aromatic polyimides is set forth in the following Equation (1):



n is 5 to 100.

There are many aromatic dianhydrides which may be employed in the preparation of linear aromatic polyimides. Some of the more commonly used aromatic dianhydrides are:

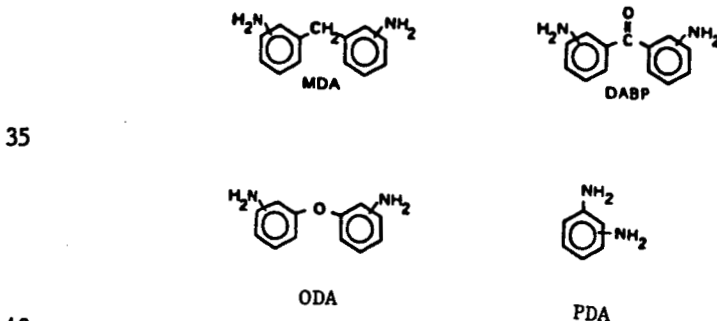
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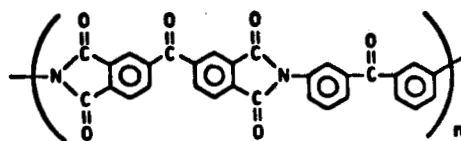
25

Likewise, there are a large member of aromatic diamines which may be employed in the preparation of linear aromatic polyimides. Some of the more commonly used aromatic diamines are:

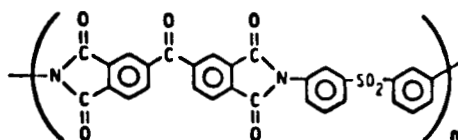
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Two particular linear aromatic polyimides--designated "LARC-TPI" and "PISO2"--will be employed herein to illustrate the process of the present invention:

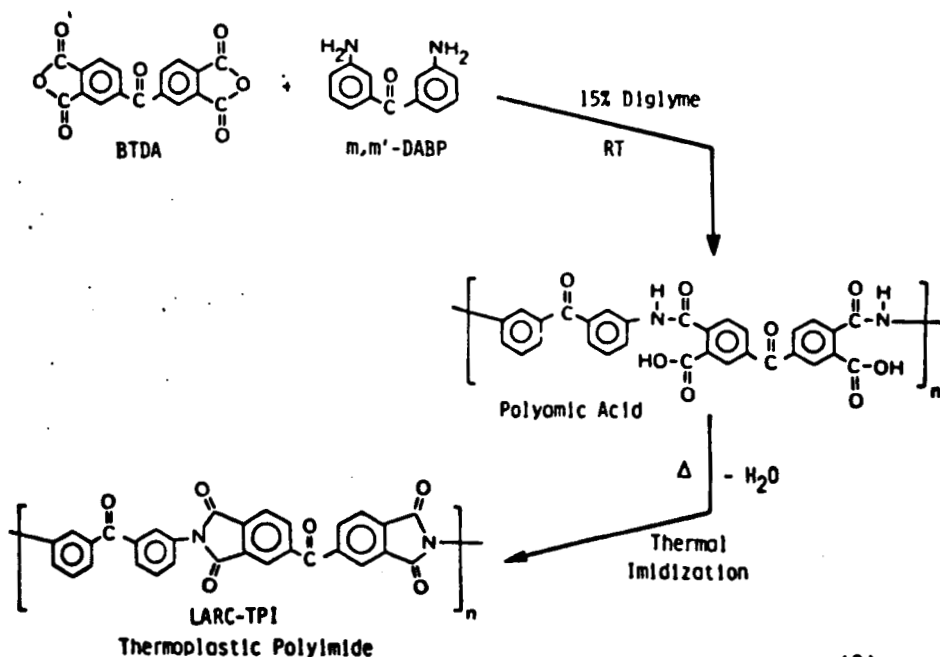


LARC-TPI



PISO2

The preparation of LARC-TPI is set forth in the following Equation (2):



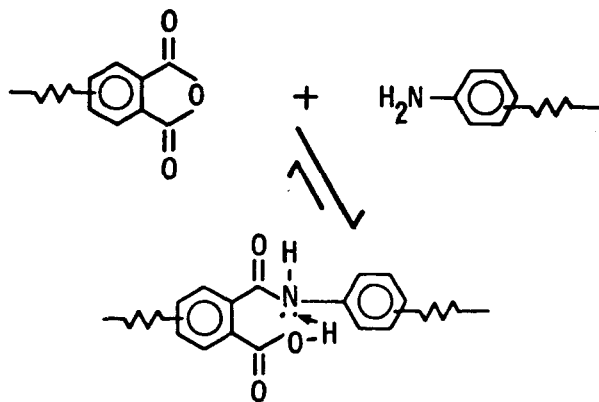
(2)

The employment of thermal dehydration in the formation of this polyimide results in a polyimide with no apparent crystallinity. An X-ray diffraction pattern for this amorphous form of LARC-TPI is shown in Figure 1. For comparison, an X-ray diffraction pattern for a semi-crystalline LARC-TPI molding powder that was manufactured by Mitsui Toatsu Fine Chemical Company (MTC) in Japan is shown in Figure 2. This pattern is illustrative of diffractions caused by crystallinity. The level of crystallinity has been estimated by MTC to be about 40%. The process by which MTC develops crystallinity in the LARC-TPI has not been disclosed. The semi-crystalline polyimide, LARC-TPI, from MTC has a melting point of about 273°C, as determined by Differential Scanning Calorimetry (DSC). However, the glass transition temperature (T<sub>g</sub>) of the resulting amorphous polyimide (after exposing the semi-crystalline polymer to a temperature of approximately 330°C, cooling to ambient temperature, and rescanning) is lower than the reported T<sub>g</sub> by about 40°C (214°C vs 254°C). This low T<sub>g</sub> is indicative of low molecular weight. In fact, when a sample of the MTC powder was dissolved in meta-cresol and an inherent viscosity (IV) was determined at 25°C at 0.5% solids using a capillary viscometer (Ostwald-Fenske type), the IV was only 0.2 dL/g, as compared to a known high molecular weight LARC-TPI which has an IV of 0.4 dL/g. The viscosity of a polymer has been shown to be directly related to the molecular weight empirically by the equation  $[\eta] = K'M^a$  where  $[\eta]$  is the intrinsic viscosity (obtained by measuring the IV at several concentrations and extrapolating to 0% concentration of solids). The IV is a good approximation of  $[\eta]$  for polyimides; K' and a are constants determined from a double logarithmic plot of intrinsic viscosity and molecular weight. The exponent, a, varies from 0.5 to 1.0, but for polyimides it is generally between 0.6 and 0.8. Typical values of K' range between 0.5 and  $5 \times 10^{-4}$ .

Both K' and a are functions of the solvent. In the case of LARC-TPI in meta-cresol, since the solvent remains the same and the constants K' and a are fixed for a

single polymer system, then the IV is an excellent approximation of molecular weight.

The explanation for the low IV for the MTC LARC-TPI is that as a polyamide-acid is cyclodehydrated to obtain the polyimide, there is a side reaction that can occur. This side reaction involves the cyclization to the anhydride with the reformation of the amine unit as shown in the following Equation (3):



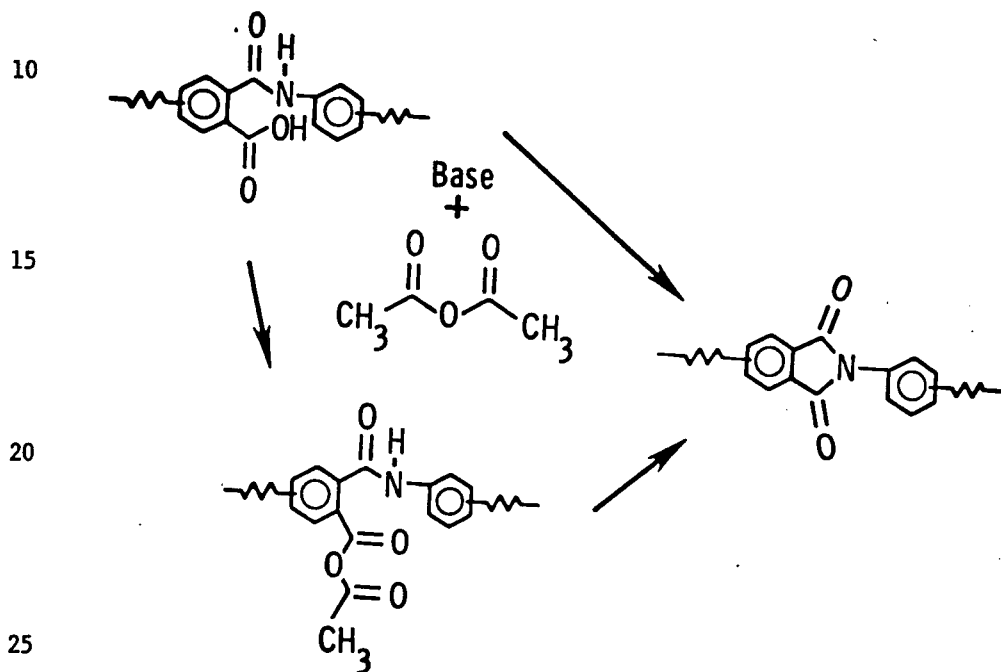
(3)

The tendency for depolymerization is influenced by the acidity of the proton, the basicity of the amide, and the nature of the solvating medium.

Chemical dehydration of polyamide-acids to form polyimides has been in use for over two decades, and is mentioned in U.S. Patents 3,179,634 and 3,179,614. However, the employment of this technique has always involved the use of a mixture of an organic base with a dehydrating agent such as acetic anhydride. The function of the organic base was always reported to be for complexing with the acetic

acid that is formed as a by-product. However, in the practice of the present invention there is evidence of an intermediate mixed anhydride as set forth in Equation (4):

5



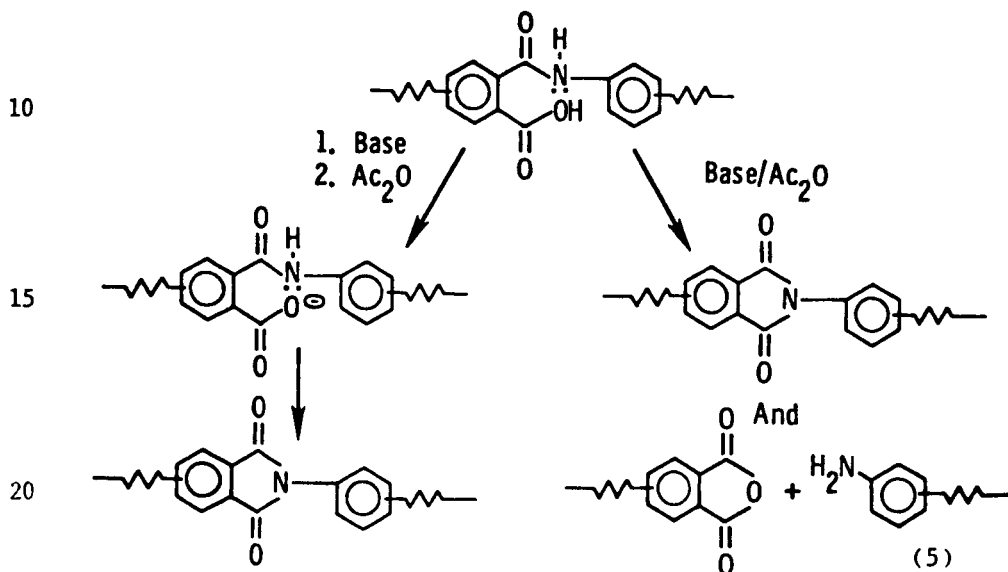
(4)

30

35

40

The formation of this intermediate can be enhanced if the organic base is added first to the polyamide-acid, as is done in the present invention. In this case the organic base can abstract a proton from the carboxylic acid moiety to form an anion as shown in the reaction scheme on the left in Equation (5):



This anion can either add an acetyl group from acetic anhydride with subsequent elimination of acetic acid as shown in Equation (4) above or simply cyclize to the imide structure directly. When this sequential addition of organic base followed by acetic anhydride is employed, there is no evidence of depolymerization as is the case where the mixture of base and anhydride is employed in one step. When polyimides are cyclized using either thermal or state-of-the-art chemical techniques, there is evidence of depolymerization from infrared spectroscopy, which shows a significant level of reformed aromatic anhydride (Young and Chang, 30th National SAMPE Symposium, Vol. 30, pp. 889-902,

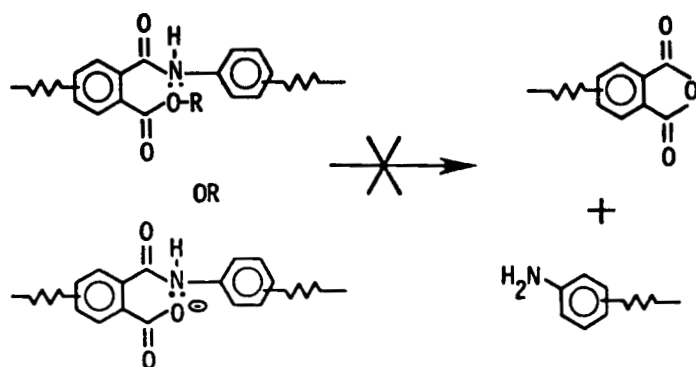


1985). Therefore, an obvious benefit of this invention is the absence of this retrograde reaction as depicted in Equation (6):

5

10

15



Where R = Acetyl

(6)

20

25

30

After diluting a 30 percent solution of LARC-TPI polyamide-acid in diglyme with N-methylpyrrolidone to yield a ten percent solution, this sequential addition of organic base followed by the acetic anhydride not only allowed for near quantitative conversion of the polyamide-acid to the polyimide, but it also allowed for the formation of a significant level of crystallinity in the resulting polyimide. Figure 3 is the X-ray diffraction pattern of LARC-TPI which was prepared using this novel technique with triethylamine as the base. The highly refined pattern indicates a level of crystallinity higher than for the MTC LARC-TPI (Figure 2).

35

For comparison, the same LARC-TPI polyamide-acid solution was treated with the same two reagents in the reverse order; that is, first the acetic anhydride was added and stirred for five minutes, and next the organic base, triethylamine, was added and stirring continued for one hour. The X-ray diffraction pattern for this polymer is shown in Figure 4. This polymer exhibits a relatively small

40

amount of crystallinity when compared to the polymer in Figure 3.

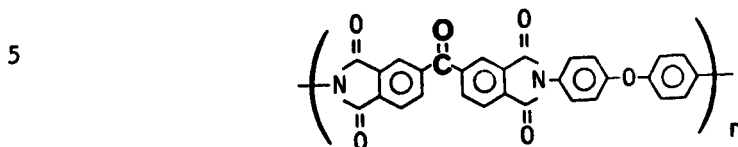
5 In order to determine if other aprotic organic bases would work, pyridine was substituted for triethylamine, and the same procedure was followed as for the polymer that yielded the X-ray diffraction pattern shown in Figure 3. The X-ray diffraction pattern for the polyimide from this pyridine-followed-by-acetic-anhydride process is shown in  
10 Figure 5. It is evident that this material is nearly identical in crystalline structure and content to the polyimide that afforded the X-ray diffraction pattern shown in Figure 3.

Figure 6 shows the X-ray diffraction pattern of the  
15 same polymer, LARC-TPI, when the polyamide-acid solution was diluted with more diglyme instead of N-methylpyrrolidone to yield a 10 percent solution. Without the presence of the amide solvent there was practically no crystallinity evident.

20 In addition, the chemically imidized LARC-TPI which afforded the crystalline X-ray diffraction shown in Figure 3 was redissolved in meta-cresol, and an inherent viscosity was determined to be 0.4 dL/g. This indicates high molecular weight (as compared to the MTC LARC-TPI which had  
25 an inherent viscosity of 0.2 dL/g).

Another polyamide-acid in diglyme at 30 percent solids, polyimidesulfone or PISO2 as structurally shown supra, was diluted with N-methylpyrrolidone to yield a 10 percent solids solution. This solution was stirred with an equal  
30 volume of triethylamine for 30 minutes. Next a volume of acetic anhydride equal to the triethylamine was added, and stirring continued for 48 hours. After the powdered polyimide was collected and dried, it was found to be crystalline by X-ray diffraction (Figure 7). A sample of  
35 PISO2 that was imidized by heating the polyamide-acid powder to 300°C was found to be totally amorphous as shown in Figure 8.

A commonly used polyimide of the following structure



where n is >5

10 and designated BTDA-ODA was chemically imidized in the same manner as was the PISO2 in the foregoing discussion. No reports of significant amounts of crystallinity have been previously issued for this polymer. However, the polyimide  
15 of BTDA-ODA made by this invention exhibited a significant level of crystallinity as evidenced by the X-ray diffraction seen in Figure 9.

The melt flow behavior of the crystalline MTC LARC-TPI was compared to that of the NASA LARC-TPI that afforded the  
20 X-ray diffraction in Figure 3 (viz., that which was made according to the subject invention). The apparent viscosity as a function of strain rate is plotted in Figure 10. Even though the NASA LARC-TPI had a higher inherent viscosity than did the MTC LARC-TPI (0.4 vs. 0.2 dL/g), the NASA  
25 LARC-TPI exhibited slightly lower viscosities than did the MTC-LARC-TPI. This difference is likely attributable to a higher level of crystallinity in the NASA LARC-TPI since the lower molecular weight MTC polymer should have exhibited lower viscosity if all other factors were equivalent.

30 The following specific examples are presented to further illustrate, but not to limit the present invention, and to reveal the best mode known to the inventor for putting the present invention into practice.

#### 35 EXAMPLE I

The polyamide-acid of LARC-TPI at 29.1% solids in diglyme (bis-2-methoxyethylether) was obtained from Mitsui Toatsu Corporation (MTC), Tokyo, Japan. This LARC-TPI  
40 (50 ml) was diluted with 100 ml of diglyme. Next was added a solution composed of 300 ml of acetic anhydride and 300 ml

of triethylamine (TEA) which had been previously mixed in a separate flask. This solution was stirred overnight, and the solid that precipitated was collected by suction  
5 filtration and was washed with 500 ml of water. Next this powder was dried in air for 30 minutes and then placed in an air oven at 100°C for two hours. This powder was analyzed by X-ray diffraction and DSC. No crystallinity was evident.

#### 10 EXAMPLE II

The same polyamide-acid as in Example I was diluted with N-methylpyrrolidone (NMP). That is, 50 ml of the MTC polyamide-acid was diluted with 100 ml of NMP. The same  
15 procedure as in Example I was followed. Only a trace of crystallinity was evident in the X-ray diffraction pattern.

#### EXAMPLE III

Another sample of the MTC polyamide-acid that had been diluted with NMP as in Example II was stirred and 300 ml of TEA was added and the mixture continued to stir for 10 minutes. Next 300 ml of acetic anhydride was added and stirring continued for one hour. The workup of this solid  
20 polyimide was the same as in Example I. In this Example the DSC indicated a melt endotherm that had a minimum at about 270°C. The X-ray diffraction exhibited a pattern typical of a crystalline structure (Figure 3). The estimate of  
25 crystallinity from the X-ray diffraction was about 50% or higher. The inherent viscosity of this polyimide was found to be 0.4 dL/g when run at 0.5% solids in meta-cresol at  
30 25°C using an Ostwald-Fenske viscometer.

#### EXAMPLE IV

Another 200 ml sample of the MTC polyamide-acid was diluted with 400 ml of NMP. To this was added 200 ml of TEA. This mixture was stirred in a blender for 20 minutes. Next 200 ml of acetic anhydride was added, and the mixture  
35

was stirred for 24 hours. The solid was collected and dried as in Example I. The resulting polyimide had a large endotherm with a minimum near 250°C. The X-ray diffraction pattern indicated the sample to be crystalline (similar to Figure 3). The fine particles from this reaction mixture were analyzed by X-ray diffraction and found to have a higher degree of crystallinity than the courser particles.

#### EXAMPLE V

A 50 ml sample of the polyimidesulfone (PISO2) in diglyme at 29% solids was precipitated from solution by the addition of 200 ml of water. The powder was collected by suction filtration. This polyamide-acid powder was dried for one hour at 100°C, one hour at 200°C, and then one hour at 300°C. This resulting polyimide powder was found to exhibit no sign of crystallinity by either X-ray diffraction or DSC.

#### EXAMPLE VI

A 50 ml sample of PISO2 in diglyme at 29% solids was diluted with 100 ml of NMP. To this was added 150 ml of TEA and the mixture was stirred for 30 minutes. Next 150 ml of acetic anhydride was added, and the mixture was stirred for 48 hours. Water (200 ml) was added to precipitate the polyimide, because the imide form of PISO2 is soluble in NMP. The solid polymer was collected by suction filtration and was air dried for seven days at ambient conditions. The polymer was then placed in an oven and heated to 100°C and held at that temperature for 5 hours, because a large amount of acetic acid evolution occurred. Next the powder was heated to 200°C and held for three hours to ensure total elimination of acetic acid. The powder was analyzed by DSC and found to have a melt endotherm at about 230°C. The X-ray diffraction pattern (Figure 7) showed the polyimide to have a significant level of crystallinity.

#### EXAMPLE VII

A 10 ml sample of the polyamide-acid designated

15 15% solids was diluted with 20 ml of NMP. To this mixture  
was added 10 ml of TEA, and stirring was effected for 30  
minutes. Next 10 ml of acetic anhydride was added, and the  
mixture was stirred overnight. This polyimide precipitated  
in large chunks, which were collected by suction filtration.  
They were air dried overnight and heated in an air oven for  
one hour at 100°C. This material was ground to a powder,  
after which the X-ray diffraction pattern thereof (Figure 9)  
10 was determined. A significant level of crystallinity was  
indicated.

#### EXAMPLE VIII

15 A 50 ml sample of the MTC LARC-TPI polyamide-acid as in  
Example II was diluted to 10 percent solids with NMP. To  
this was added 150 ml of pyridine, and the reaction mixture  
was stirred for 30 minutes. Next was added 150 ml of acetic  
anhydride, and the mixture was stirred for one hour and  
20 allowed to stand without stirring for 48 hours. The solid  
was collected by suction filtration and was washed with 500  
ml of water. The solid was air dried for one hour and then  
heated to 100°C in an air oven for one hour. The X-ray  
diffraction pattern showed the polyimide to be crystalline  
25 (Figure 5).

According to the present invention, polyamide-acids are  
chemically dehydrated to form polyimides which are  
semi-crystalline and which have melting temperatures near  
30 the glass transition temperatures of the amorphous forms of  
the polyimides. Upon melting of these semi-crystalline  
polyimides, the crystalline forms do not reform. Flow is  
thereby enhanced as melting occurs. Moreover, the  
semi-crystalline polyimides are prepared without degradation  
35 of the molecular weight of the polyimides.

The present invention has been described in detail with  
respect to certain preferred embodiments thereof. As is  
understood by those of skill in the art, variations and  
modifications in this detail may be effected without any  
40 departure from the spirit and scope of the present  
invention, as defined in the hereto-appended claims.

PROCESS FOR DEVELOPING CRYSTALLINITY IN LINEAR AROMATIC  
POLYIMIDES5    Abstract of the Invention

          The process of the present invention includes first  
treating a polyamide-acid (such as LARC-TPI polyamide-acid)  
in an amide-containing solvent (such as N-methylpyrrolidone)  
10 with an aprotic organic base (such as triethylamine),  
followed by dehydrating with an organic dehydrating agent  
(such as acetic anhydride).

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